National
Qualifications
2026

X813/76/12

Chemistry
Paper 1 — Multiple choice

TUESDAY, 12 MAY
9:00 AM – 9:40 AM

Total marks — 25

Attempt ALL questions.

You may use a calculator.

Instructions for the completion of Paper 1 are given on *page 02* of your answer booklet X813/76/02.

Record your answers on the answer grid on *page 03* of your answer booklet.

You may refer to the Chemistry Data Booklet for Higher and Advanced Higher.

Space for rough work is provided at the end of this booklet.

You must leave your answer booklet on your desk; if you do not, you could lose all the marks for this paper.

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Total marks — 25

Attempt ALL questions

1. The bonding and structure in fullerenes is

A ionic lattice
B metallic lattice
C covalent network
D covalent molecular.

2. The two hydrogen atoms in a molecule of hydrogen are held together by a

A hydrogen bond
B polar covalent bond
C pure covalent bond
D van der Waals force.

3. Which of the following statements is true?

A The oxygen atom is larger than the oxide ion.
B The magnesium atom is larger than the magnesium ion.
C The sulfide ion is smaller than the sulfur atom.
D The calcium ion is larger than the calcium atom.

4. What is the enthalpy change, in $kJ\,mol^{-1}$, for the following reaction?

$$Sc(g) \rightarrow Sc^{2+}(g) + 2e^{-}$$

A 633
B 1235
C 1868
D 2470

5. Which of the following is a redox reaction?

A $Zn + 2HCl \rightarrow ZnCl_2 + H_2$
B $NaOH + HCl \rightarrow NaCl + H_2O$
C $NiO + 2HCl \rightarrow NiCl_2 + H_2O$
D $CuCO_3 + 2HCl \rightarrow CuCl_2 + H_2O + CO_2$

6. Iodide ions can be oxidised to iodine.

$$2I^-(aq) \rightarrow I_2(s) + 2e^-$$

Which of the following ions could be used to oxidise iodide ions to iodine?

A $SO_3^{2-}(aq)$

B $SO_4^{2-}(aq)$

C $Cr^{3+}(aq)$

D $Cr_2O_7^{2-}(aq)$

7. Oil molecules are more likely to react with oxygen in the air than fat molecules.
During the reaction the oil molecules

A are reduced

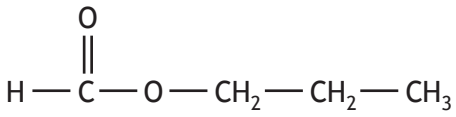
B form rancid products

C are hydrolysed

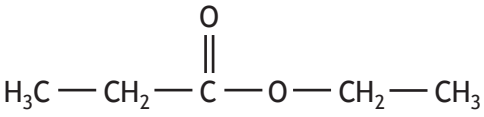
D become unsaturated.

8. A student was given samples of ethanol, propan-1-ol, methanoic acid and ethanoic acid.
Which of the following esters **could not** be made from these chemicals?

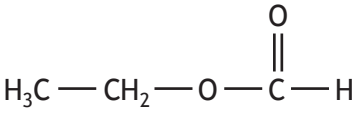
A



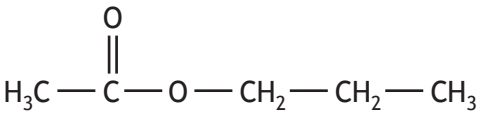
B



C



D



[Turn over

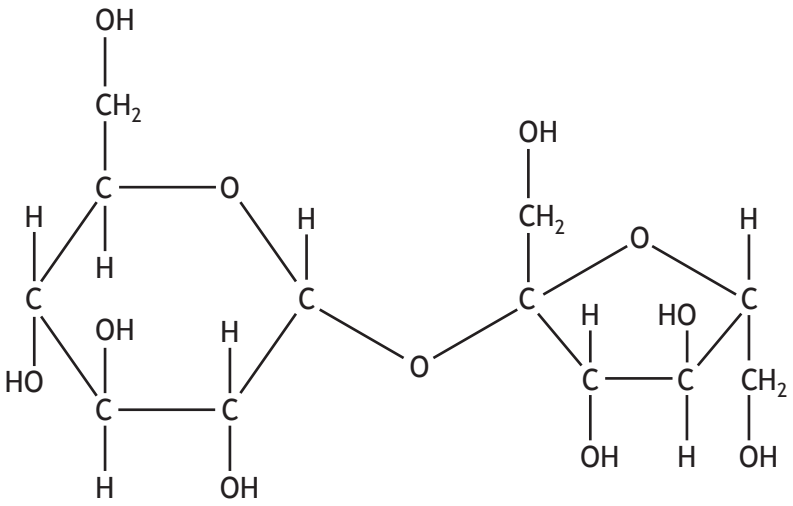
9. A compound with the formula CH_2O_2 must be

A an ester
B an alcohol
C an aldehyde
D a carboxylic acid.

10. Which process is used to convert ethanoic acid to ethanal?

A Addition
B Condensation
C Hydrolysis
D Reduction

11. Olestra is a calorie free fat made by reacting carboxylic acids with sucrose.
The structure of a sucrose molecule can be represented as shown.



How many carboxylic acid molecules can react with one molecule of sucrose?

A 3
B 5
C 8
D 11

12. Fats have higher melting points than oils because

A fats are more saturated
B oils are more saturated
C fats are less closely packed
D oils are more closely packed.

13. Propanal, C_3H_6O, can be converted to propanoic acid by heating with acidified dichromate solution.

$$Cr_2O_7^{2-} + 14H^+ + 6e^- \rightarrow 2Cr^{3+} + 7H_2O$$

$$C_3H_6O + H_2O \rightarrow C_3H_6O_2 + 2H^+ + 2e^-$$

How many moles of $Cr_2O_7^{2-}$ would react with one mole of propanal?

A $\frac{1}{3}$

B $\frac{2}{3}$

C 1

D 3

14. Which of the following is an isomer of 2,2-dimethylpentan-1-ol?

A $CH_3CH_2CH_2C(CH_3)_2CH_2OH$

B $(CH_3)_3CCH(CH_3)CH_2OH$

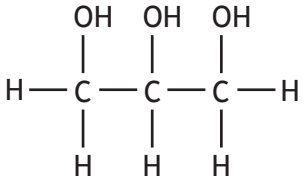
C $CH_3CH_2CH_2CH_2CH_2CH_2CH_2CH_2OH$

D $(CH_3)_2CHC(CH_3)_2CH_2CH_2OH$

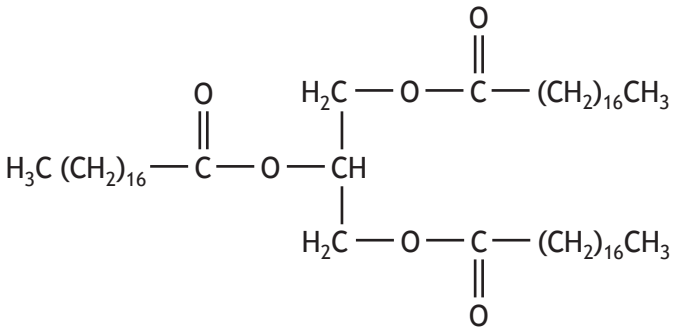
[Turn over

15. Which of the following could be made by reacting edible oils with glycerol?

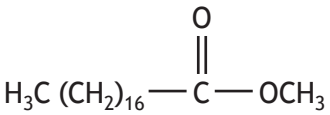
A



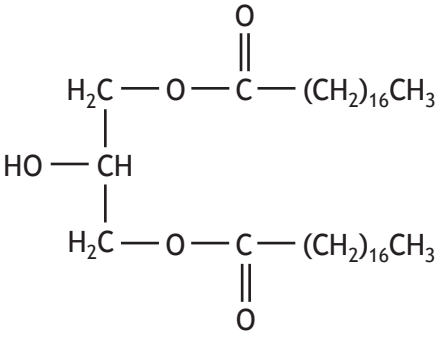
B



C



D



16. A two-step reaction is shown.

Step 1

$$S_2O_8^{2-}(aq) + 2I^-(aq) + 2Fe^{2+}(aq) \rightarrow 2SO_4^{2-}(aq) + 2I^-(aq) + 2Fe^{3+}(aq)$$

Step 2

$$2SO_4^{2-}(aq) + 2I^-(aq) + 2Fe^{3+}(aq) \rightarrow 2SO_4^{2-}(aq) + I_2(aq) + 2Fe^{2+}(aq)$$

The ion that is the catalyst in the overall reaction is

A $S_2O_8^{2-}(aq)$

B $I^-(aq)$

C $Fe^{2+}(aq)$

D $SO_4^{2-}(aq)$

17. A student added 0.1 mol of zinc to a solution containing 0.05 mol of silver(I) nitrate.

$$Zn(s) + 2AgNO_3(aq) \rightarrow Zn(NO_3)_2(aq) + 2Ag(s)$$

Which of the following statements about the experiment is correct?

A 0.05 mol of zinc reacts.

B 0.05 mol of silver is produced.

C Silver nitrate is in excess.

D All of the zinc reacts.

[Turn over

18. The following equilibrium exists as an orange solution.

$$\underset{\text{(yellow)}}{Fe^{3+}(aq)} + CNS^{-}(aq) \rightleftharpoons \underset{\text{(red)}}{[FeCNS]^{2+}(aq)}$$

Addition of **X** causes the reaction mixture to become red.

X could be

A $FeCl_3$

B $[FeCNS]Cl_2$

C NaCl

D $FeCl_2$

19. Which of the following is **not** a factor that affects the rate of a reaction?

A Activation energy

B Kinetic energies of reactant molecules

C Concentration of reactants

D Enthalpy change of reaction

20. Chromatography separates substances by making use of their differences in size and

A boiling point

B polarity

C reactivity

D colour.

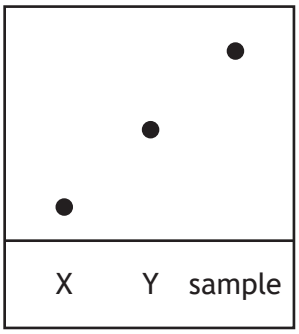
21. A chemist is attempting to make a fragrance compound by the following chemical reaction.

compound **X** + compound **Y** → fragrance compound

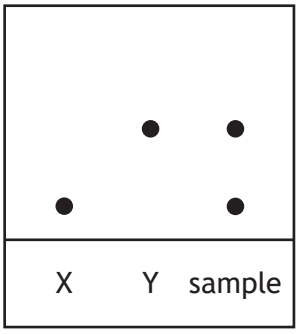
Compounds **X** and **Y** are mixed in a flask. After some time, a sample is removed from the flask and compared with pure samples of compounds **X** and **Y** using chromatography.

Which of the following chromatograms shows that a pure sample of the fragrance compound has been produced?

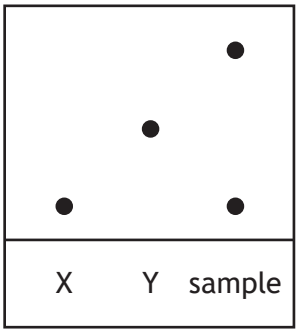
A



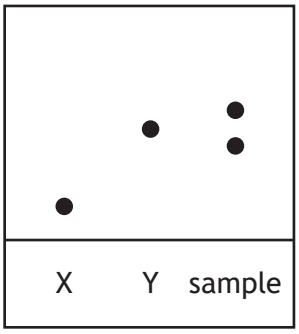
B



C



D



[Turn over

22. The vitamin C content of a carton of orange juice was determined by four students. Each student carried out the experiment three times.

	Concentration of vitamin C (mg/100 cm^3)		
	Experiment 1	**Experiment 2**	**Experiment 3**
Student A	28.0	24.0	23.0
Student B	21.4	21.6	21.8
Student C	21.9	22.0	21.9
Student D	22.0	21.5	21.9

The most reproducible results were obtained by

A student A

B student B

C student C

D student D.

23. Which of the following is the most appropriate apparatus to accurately prepare a 0.10 mol l^{-1} solution from a 1.0 mol l^{-1} solution?

A 1 cm^3 dropper and a 10 cm^3 measuring cylinder.

B 10 cm^3 volumetric flask and a 100 cm^3 conical flask.

C 25 cm^3 pipette and a 250 cm^3 volumetric flask.

D 50 cm^3 burette and a 500 cm^3 measuring cylinder.

24. The octane number is a measure of how easily hydrocarbons ignite.

	Hydrocarbon	Octane number
1	3-methylpentane	74.5
2	pentane	61.7
3	butane	93.6
4	2-methylpentane	73.4
5	hexane	24.8
6	methylcyclopentane	91.3

A student predicted that as the chain length of a hydrocarbon increases, the octane number decreases.

Which set of three hydrocarbons should have their octane numbers compared to provide the best test of this prediction?

A 1, 4, 6

B 1, 2, 4

C 3, 4, 5

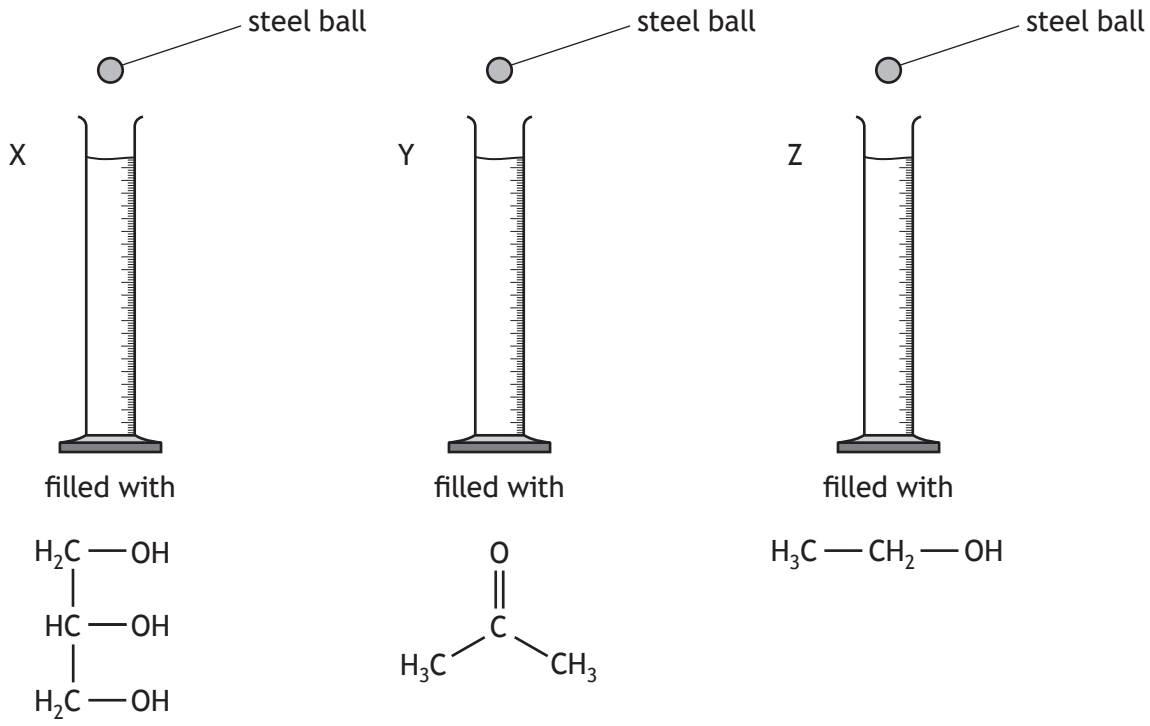
D 2, 3, 5

[Turn over

25. The type of forces between molecules can affect the viscosity of a liquid.

A steel ball will fall at different speeds through different liquids.

The more viscous a liquid the more slowly the ball falls through the liquid.



Which of the following shows the correct order for the speed of the steel balls falling through the liquids from the fastest to the slowest?

A Y, X then Z

B X, Z then Y

C Y, Z then X

D Z, Y then X

[END OF QUESTION PAPER]

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National Qualifications 2026

Mark

X813/76/01

Chemistry Paper 2

TUESDAY, 12 MAY

10:10 AM – 12:30 PM

X8137601

Fill in these boxes and read what is printed below.

Full name of centre

Town

Forename(s)

Surname

Number of seat

Date of birth

Day | Month | Year

Scottish candidate number

Total marks — 95

Attempt ALL questions.

You may use a calculator.

You may refer to the Chemistry Data Booklet for Higher and Advanced Higher.

Write your answers clearly in the spaces provided in this booklet. Additional space for answers and rough work is provided at the end of this booklet. If you use this space you must clearly identify the question number you are attempting. Any rough work must be written in this booklet. Score through your rough work when you have written your final copy.

Use **blue** or **black** ink.

Do not remove any exam materials. You must leave this booklet on your desk; if you do not, you could lose all the marks for this paper.

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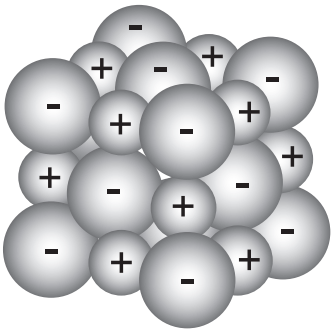
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Total marks — 95

Attempt ALL questions

1. (a) Sodium chloride is an ionic compound and has the structure shown.



(i) State the term used to describe the structure of sodium chloride. **1**

(ii) An ionic bond is a type of electrostatic attraction.

(A) Describe what is meant by the term electrostatic attraction. **1**

(B) Explain why sodium and chlorine form an ionic compound using electronegativity values from the data booklet. **1**

1. (continued)

(b) The boiling point of pure water is 100 °C. Dissolving an ionic solute in water raises the boiling point.

The increase in boiling point depends on the total number of solute particles.

The increase in boiling point (ΔT_b), in °C, can be calculated using the formula shown.

$$\Delta T_b = 0.51 \times c \times i$$

c = the concentration of the solution in mol l^{-1}

i = the number of particles in solution when one formula unit of solute dissolves

The values of i for some ionic solutes are shown in the table.

Ionic solute	Value of i
KCl	2
$CaBr_2$	3
$(NH_4)_3PO_4$	4

(i) State the value of **i** for lithium sulfate. **1**

(ii) Calculate the boiling point, in °C, of an aqueous solution of 0.35 mol l^{-1} of ammonium phosphate. **2**

[Turn over

2. Water, H_2O, and methane, CH_4, are covalent molecules.

The table shows the van der Waals forces present and boiling points for these molecules.

Molecule	Van der Waals forces present	Boiling point (°C)
CH_4	London dispersion forces	−162
H_2O	London dispersion forces Hydrogen bonds	100

(a) Explain **fully** how London dispersion forces are formed. **2**

(b) The boiling points of these two molecules, H_2O and CH_4, can be used to compare the strength of van der Waals forces.

Explain why H_2O and CH_4 are suitable for this comparison. **1**

(c) Explain **fully** why water is a polar molecule. **2**

3. Trends in melting points, ionisation energy, covalent radius and electronegativity can be related to the position of an element in the periodic table.

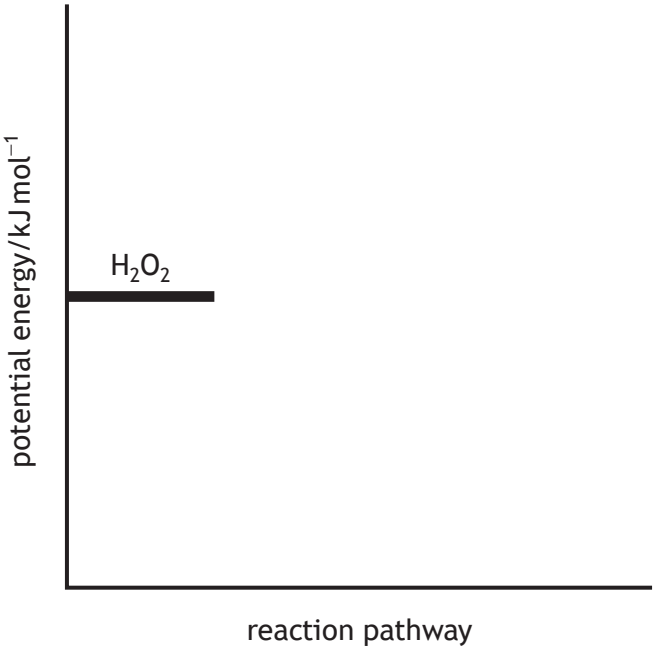
Using your knowledge of chemistry, comment on how the position of an element in the periodic table can be used to predict its physical and chemical properties. **3**

[Turn over

4. Hydrogen peroxide, H_2O_2, decomposes very slowly to produce water and oxygen.

$$2H_2O_2(aq) \rightarrow 2H_2O(\ell) + O_2(g)$$

(a) The activation energy, E_a, for this reaction is 75 kJ mol^{-1} and the enthalpy change, ΔH, is -98 kJ mol^{-1}.



(i) Use the activation energy and enthalpy change given above to complete the potential energy diagram for the reaction. **1**

(An additional diagram, if required, can be found on *page 32*.)

(ii) Calculate the activation energy, in kJ mol^{-1}, for the **reverse** reaction. **1**

4. (continued)

(b) Manganese dioxide catalyses the decomposition of hydrogen peroxide solution.

When a catalyst is used in a reaction more successful collisions occur and the rate of reaction increases.

Explain **fully** the way in which catalysts increase the number of successful collisions. **2**

(c) The balanced equation for the decomposition of hydrogen peroxide, H_2O_2, is shown.

$$2H_2O_2(aq) \rightarrow 2H_2O(\ell) + O_2(g)$$

$$GFM = 34 \text{ g}$$

Calculate the mass, in g, of hydrogen peroxide required to produce 60 cm^3 of oxygen gas. **2**

Take the volume of one mole of oxygen gas to be 24 litres.

[Turn over

4. (continued)

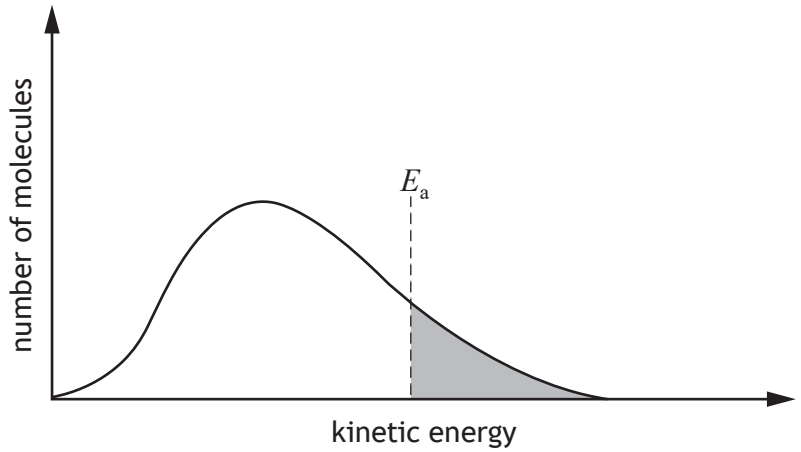
(d) Hydrogen peroxide, H_2O_2, should be stored in the dark as sunlight can provide sufficient energy to break the weak oxygen to oxygen single bond in the molecule creating free radicals.

(i) State the meaning of the term free radicals. **1**

(ii) Identify the free radical formed from hydrogen peroxide. **1**

(e) Increasing the temperature will increase the rate of a reaction.

The diagram shows the distribution of kinetic energies of molecules in a sample of a gas.



(i) Suggest what is represented by the shaded area in the diagram. **1**

(ii) Add a second curve to the diagram to show the distribution of kinetic energies at a higher temperature. **1**

(An additional diagram, if required, can be found on *page 32*.)

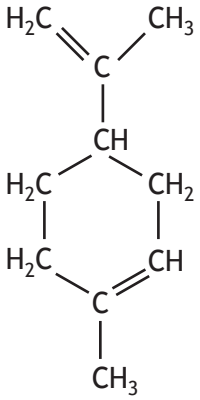
5. The peel of citrus fruits contains essential oils.

(a) Essential oils in citrus fruits have aromas.

State another property of an essential oil. **1**

(b) Limonene is an essential oil found in citrus peel.

A structure for limonene is shown.



(i) When limonene is heated, it can undergo decomposition to form isoprene units.

Draw a structural formula for isoprene. **1**

(ii) State the number of isoprene units that would be produced when one molecule of limonene decomposes. **1**

[Turn over

5. (continued)

(c) Citrus peel has been investigated as a possible source of biofuel for cars.

(i) The first step in producing the fuel is to extract essential oils, including limonene, from citrus peel.

One method uses hexane to dissolve the essential oils.

Explain why hexane is a suitable solvent for essential oils. **1**

(ii) An experiment was carried out to determine the enthalpy of combustion for limonene.

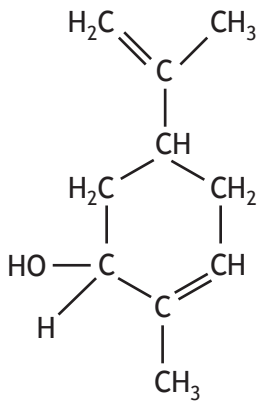
3.5 g of limonene ($GFM = 136$ g) was burned and the temperature rise in 500 cm^3 of water was 56 °C.

Calculate the enthalpy of combustion, in $kJ\,mol^{-1}$, of limonene using these experimental results. **3**

5. (continued)

(d) Limonene can undergo oxidation to produce carveol.

A structure for carveol is shown.



(i) Explain why carveol can be classified as a secondary alcohol. **1**

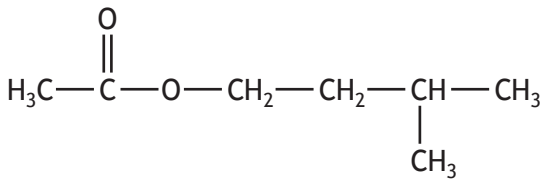
(ii) Carveol can be further oxidised to produce a ketone.

Draw a structural formula for the ketone produced when carveol is oxidised. **1**

[Turn over

6. Bananas contain a range of compounds.

(a) An ester responsible for the flavour of bananas is shown.



(i) Name the alcohol used to make this ester. **1**

(ii) This alcohol was passed over heated black copper(II) oxide.
Suggest what would be seen at the end of the reaction. **1**

6. (continued)

(b) Bananas produce the gas ethene, C_2H_4, ($GFM = 28$ g) which causes them to ripen.

To speed up the ripening of bananas, they are stored in a gas mixture containing 4% ethene and 96% nitrogen.

A gas mixture has a volume of 27 000 litres.

Calculate the mass of ethene required, in g, for this gas mixture. **2**

Take the volume of one mole of ethene gas to be 23 litres.

(c) When bananas are cut, the colour changes from yellow to brown.

Enzymes cause the breakdown of chemicals in the banana, resulting in this colour change.

(i) Name the class of compounds to which enzymes belong. **1**

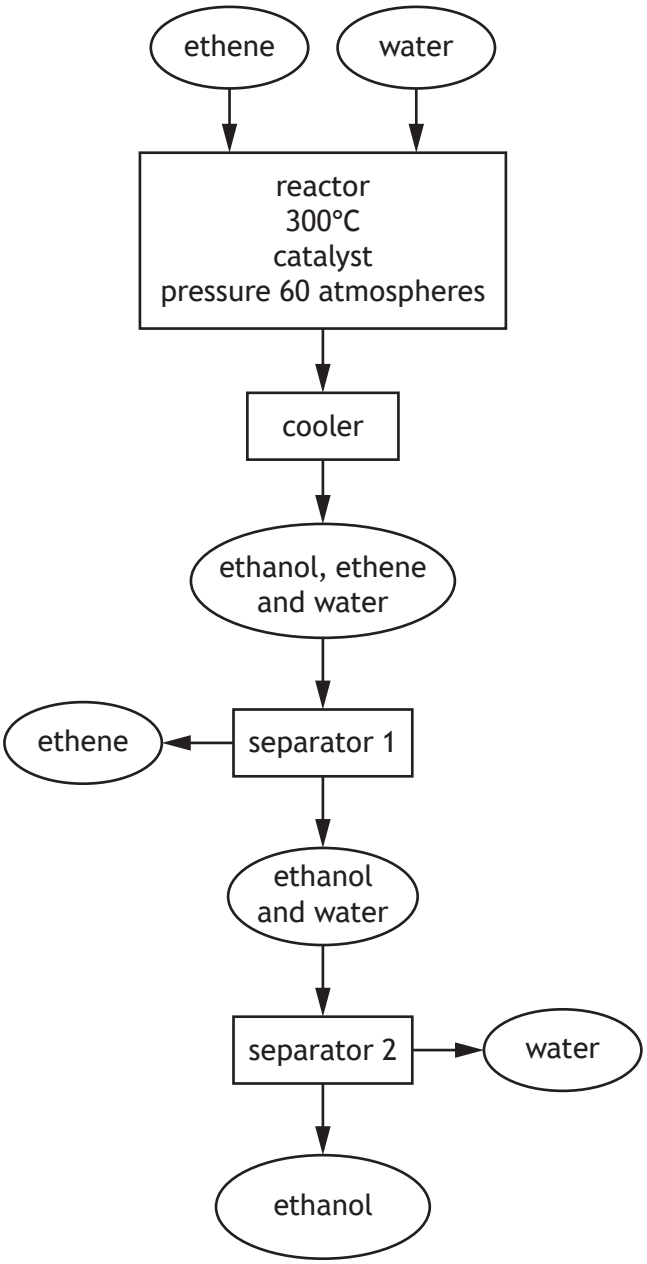
(ii) To stop cut bananas turning brown, they can be treated with lemon juice, denaturing the enzymes.

Explain **fully** what is meant by denaturing. **2**

[Turn over

7. Ethanol is an important industrial chemical.

(a) The flow diagram of an industrial process to produce ethanol is shown.



(i) **On the flow diagram**, draw an arrow to suggest one way to maximise profit in this industrial process. **1**

(ii) Name the process used in separator 2 to separate ethanol from water. **1**

7. (a) (continued)

MARKS DO NOT WRITE IN THIS MARGIN

(iii) In the reactor ethene reacts with water to form ethanol.

$$C_2H_4(g) + H_2O(g) \rightleftharpoons C_2H_5OH(g)$$

(A) The reaction taking place is called hydration.

State another name for this type of reaction. **1**

(B) Increasing the pressure in the reactor improves the yield of ethanol.

Explain **fully** why increasing pressure would shift the equilibrium to the right producing more ethanol. **2**

(C) Using mean bond enthalpies from the data booklet, calculate the enthalpy change, in kJ mol^{-1}, for this reaction. **2**

$$C_2H_4(g) + H_2O(g) \rightleftharpoons C_2H_5OH(g)$$

(D) The atom economy for this reaction is 100%.

Explain what this means. **1**

[Turn over

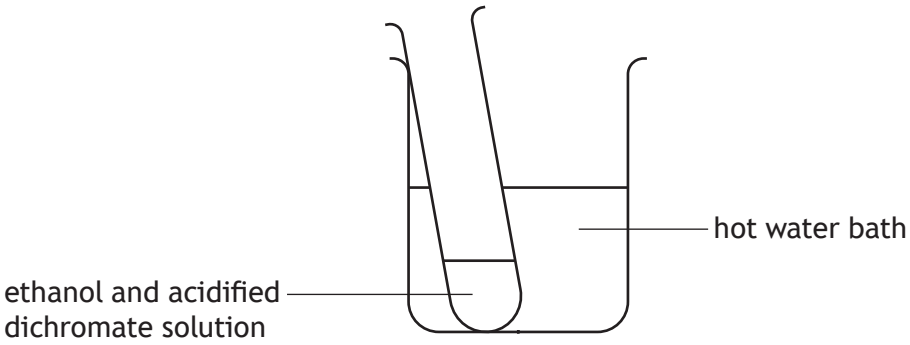
(b) Ethanol can be oxidised to compound **X**.

$$\underset{\text{ethanol}}{CH_3CH_2OH} \xrightarrow{\text{Step 1}} \underset{\text{ethanal}}{CH_3CHO} \xrightarrow{\text{Step 2}} \underset{\text{compound X}}{CH_3COOH}$$

(i) State why steps 1 and 2 can be described as oxidation. **1**

(ii) Name compound **X**. **1**

(iii) In the laboratory, acidified dichromate solution can be used to convert ethanol to compound **X**.



(A) Suggest why a water bath is a suitable method of heating the reaction mixture. **1**

(B) Describe the colour change that would be observed. **1**

7. (b) (iii) (continued)

(C) Complete the ion electron equation for the reaction. **1**

$C_2H_5OH \rightarrow CH_3COOH$

(D) The other reaction taking place is shown.

$Cr_2O_7^{2-}(aq) + 14H^+(aq) + 6e^- \rightarrow 2Cr^{3+}(aq) + 7H_2O(\ell)$

Explain why the dichromate solution needs to be acidified. **1**

(E) State the classification of alcohols that do not react with acidified dichromate solution. **1**

(iv) Name another oxidising agent that could be used to convert ethanal to compound **X** in step 2. **1**

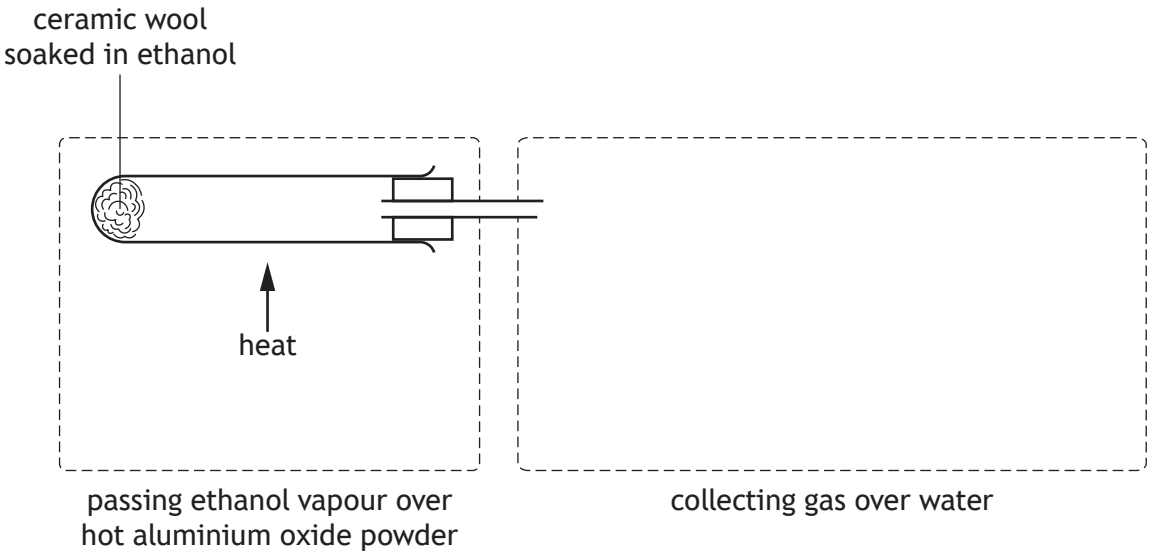
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7. (continued)

(c) Ethanol can be converted to ethene in the laboratory using a catalyst.

In this experiment, aluminium oxide powder is heated. This heat also causes the ethanol to turn to vapour.

The ethanol vapour passes over the hot aluminium oxide powder and forms ethene. The ethene gas is collected over water.



(i) Complete both parts of the diagram to show how ethanol vapour is passed over hot aluminium oxide powder and how the ethene gas is collected over water.

Your answer should include labels. **2**

(An additional diagram, if required, can be found on *page 33*.)

(ii) When bromine solution is added to the collected gas it rapidly decolourises.

State the term given to a chemical that rapidly decolourises bromine solution. **1**

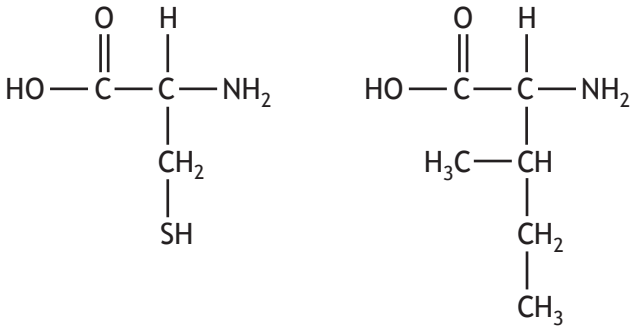
8. Snake venom contains proteins.

(a) Proteins are formed by joining together amino acids.

(i) Name the type of reaction that takes place when amino acids are joined. **1**

(ii) Two amino acids can join to form a dipeptide.

Two of the amino acids found in snake venom are shown.



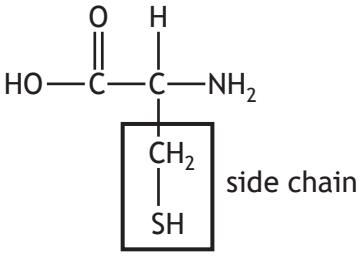
Draw a structural formula to show a dipeptide formed when these two amino acids join. **1**

[Turn over

8. (continued)

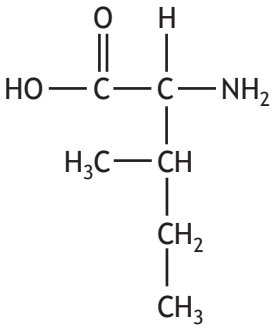
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(b) Amino acids can be classified according to their side chains.



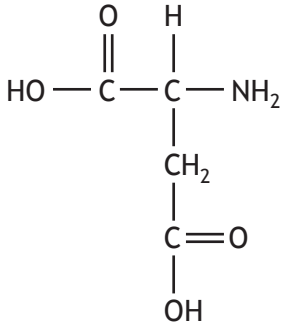
(i) Four types of side chain are **acidic**, **basic**, **hydrophobic** and **polar**.

(A) The structure of the amino acid isoleucine is shown.



Identify the class of amino acid to which isoleucine belongs. **1**

(B) The structure of the amino acid aspartic acid is shown.



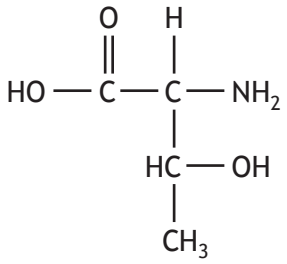
Suggest why aspartic acid can be classified as acidic. **1**

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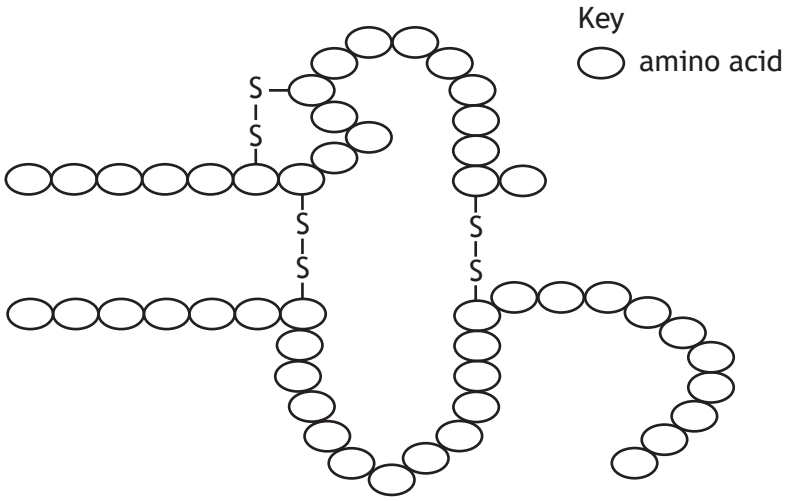
(ii) When amino acids form proteins, the proteins are stabilised by intermolecular forces that form between the side chains.

The structure of the amino acid threonine is shown.



Identify the strongest intermolecular force that would form between molecules of threonine. **1**

(c) Proteins that contain the amino acid cysteine can form disulfide bonds (— S — S —) between cysteine side chains. This makes the protein resistant to being broken down when heated.



Suggest why proteins forming disulfide bonds are resistant to heat. **1**

[Turn over

9. Liquids and powders that remove dirt and stains from clothes contain a variety of chemicals.

The chemicals in the liquids and powders include soaps, detergents, bleaching agents and enzymes.

Using your knowledge of chemistry, explain the action of the chemicals used to remove dirt and stains from clothes. **3**

10. Ammonia, NH_3, can be made using the Haber process.

(a) In the Haber process, nitrogen and hydrogen react together in the presence of an iron catalyst.

$$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g) \qquad \Delta H = -92 \text{ kJ mol}^{-1}$$

(i) Complete the table to show the change in conditions required to increase the yield of ammonia. **1**

Condition	Increase/Decrease
Pressure	increase
Temperature	
Concentration of ammonia in the reaction mixture	

(ii) State the effect of the catalyst on the equilibrium position of this reaction. **1**

(iii) Under certain conditions, 500 kg of nitrogen reacts with excess hydrogen to produce 405 kg of ammonia.

$$\underset{GFM = 28 \text{ g}}{N_2(g)} + 3H_2(g) \rightleftharpoons \underset{GFM = 17 \text{ g}}{2NH_3(g)}$$

Calculate the percentage yield of ammonia under these conditions. **2**

[Turn over

10. (continued)

(b) The hydrogen needed for ammonia production by the Haber process can be made by reacting methane with steam.

$$CH_4(g) + H_2O(g) \rightarrow CO(g) + 3H_2(g)$$

Calculate the enthalpy change, in kJ mol^{-1}, for this reaction using the following information. **2**

$$CO(g) + \tfrac{1}{2}O_2(g) \rightarrow CO_2(g) \qquad \Delta H = -283 \text{ kJ mol}^{-1}$$

$$H_2(g) + \tfrac{1}{2}O_2(g) \rightarrow H_2O(g) \qquad \Delta H = -242 \text{ kJ mol}^{-1}$$

$$CH_4(g) + 2O_2(g) \rightarrow CO_2(g) + 2H_2O(g) \qquad \Delta H = -803 \text{ kJ mol}^{-1}$$

10. (continued)

(c) A new method for producing ammonia has recently been developed.

Lithium oxide is reduced to lithium using a laser. Lithium then reacts with nitrogen in the air to form lithium nitride.

(i) Lithium nitride, Li_3N, then reacts with water.

$$Li_3N(s) + 3H_2O(\ell) \rightarrow NH_3(g) + 3LiOH(aq)$$

$GFM = 34.7$ g $\quad GFM = 18.0$ g $\quad GFM = 17.0$ g $\quad GFM = 23.9$ g

Calculate the atom economy for the production of ammonia. **2**

(ii) Using this method, 0.0013 g of ammonia ($GFM = 17.0$ g) was produced in 78 seconds.

Calculate the time, in hours, required to produce 1 mole of ammonia. **2**

[Turn over

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11. The concentration of a sodium hydroxide solution can be determined by titration with oxalic acid.

(a) A student made an oxalic acid solution of accurately known concentration.

(i) State the term used to describe a solution of accurately known concentration. **1**

(ii) Calculate the mass, in g, of oxalic acid ($GFM = 90$ g) needed to make 250 cm^3 of 0.050 mol l^{-1} oxalic acid solution. **1**

(iii) A section of the student's notebook is shown with the experimental procedure they used to make the solution.

1. The sample of oxalic acid was dissolved in 25 cm^3 of deionised water in the beaker.
2. The solution was poured into the 250 cm^3 volumetric flask using a funnel.
3. The solution was made up to the graduated mark with tap water so that the top of the meniscus meets the line.
4. The flask was stoppered and then inverted several times to mix the solution.

The solution prepared by the student did not have an accurately known concentration.

Suggest **two** improvements the student could make to the procedure. **2**

11. (continued)

(b) The student titrated the sodium hydroxide solution with a new oxalic acid solution using an indicator.

(i) State why it was necessary to add an indicator to this titration. **1**

(ii) The burette was clamped vertically, rinsed with the new oxalic acid solution and filled up to above 0.0 cm^3 of the scale.

Suggest one other step required before the burette is ready to begin the titration. **1**

[Turn over

MARKS

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(iii) Three 25.0 cm^3 samples of sodium hydroxide, NaOH, were titrated with a solution of 0.0517 mol l^{-1} oxalic acid, $(COOH)_2$.

The results are shown in the table.

Sample	Volume of oxalic acid used (cm^3)
1	27.4
2	27.0
3	26.9

(A) State why the volume of oxalic acid used in the calculation was taken to be 26.95 cm^3, although this is not the average of the three titre volumes in the table. **1**

(B) Calculate the concentration, in mol l^{-1}, of the sodium hydroxide solution. **3**

$$(COOH)_2(aq) + 2NaOH(aq) \rightarrow 2H_2O(\ell) + (COONa)_2(aq)$$

(C) The sodium salt of oxalic acid produced in the reaction has the formula $(COONa)_2$.

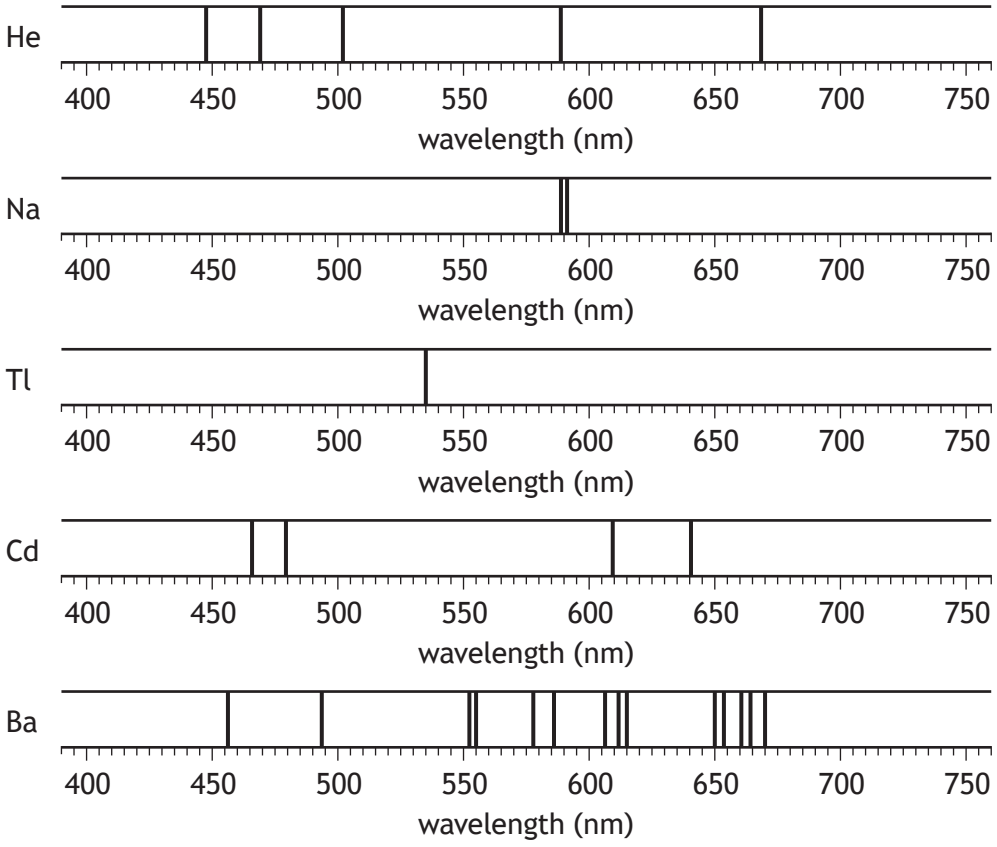
Draw a structural formula for the sodium salt of oxalic acid. **1**

12. Flame emission spectroscopy is a method of chemical analysis used to identify the elements present in a sample and determine their concentrations.

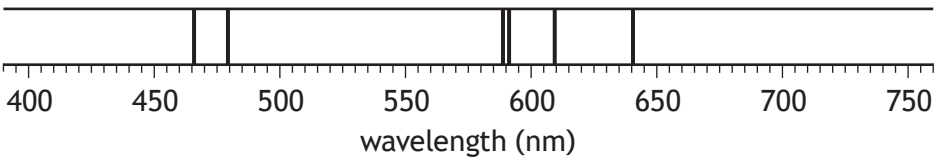
The sample is vaporised in a flame. Heat energy from the flame excites electrons in atoms to higher energy levels. When the excited electrons fall back to lower energy levels, light is emitted which is detected by the spectrometer.

Each element emits light as a pattern of characteristic wavelengths called an emission spectrum.

(a) The emission spectra for some elements are shown.



(i) A sample of water was analysed and the following emission spectrum was observed.



Name the elements present in this water sample. **1**

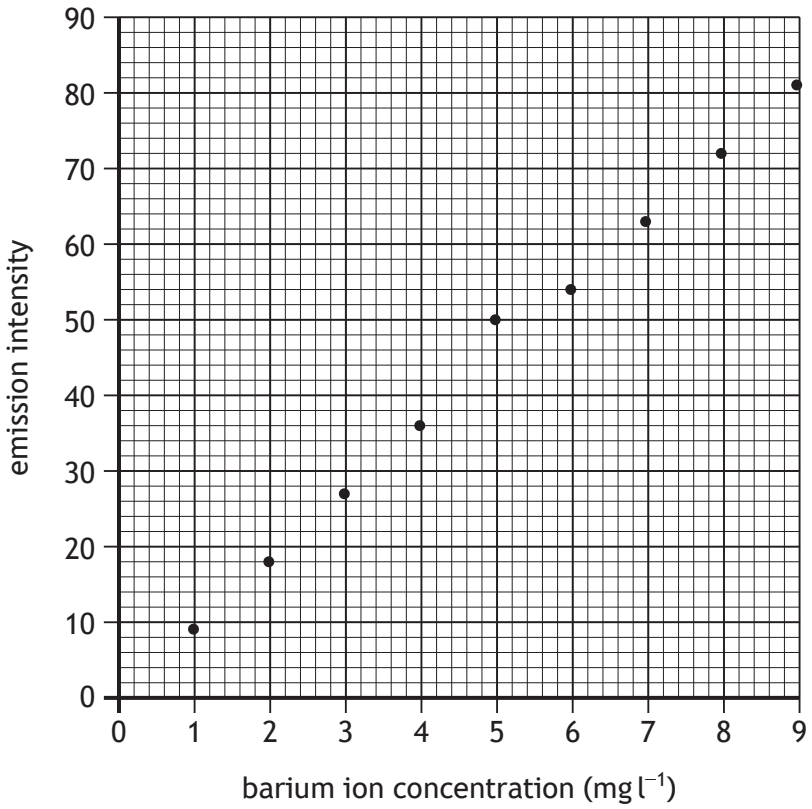
12. (a) (continued)

(ii) Thallium, Tl, shows only one line in this emission spectrum at 535 nm.

Use the colour wheel on *page 19* of the data booklet to suggest what colour flame thallium would produce. **1**

(b) The concentration of barium ions in a water sample can be determined using flame emission spectroscopy.

(i) The emission intensity was measured for a range of solutions with known concentrations of barium ions and a graph was produced.



Using the same procedure, the emission intensity of a sample was measured as 40.

Determine the concentration of barium ions, in mg l^{-1}, in the sample. **1**

12. (b) (continued)

(ii) State the term used to describe water containing high levels of dissolved metal ions. **1**

(c) The temperature of the flame is important. If it is too high, the element will ionise.

(i) Write the equation for the first ionisation energy of helium. **1**

(ii) There is no third ionisation energy value listed for helium in the data booklet.

Explain why this is the case. **1**

[END OF QUESTION PAPER]